Jumpstart #1-G

$Ca(OH)_2 + 2KBr \leftrightarrow 2KOH + CaBr_2$

1) Which way does equilibrium shift if you add extra KBr
2) What happens to the amount of KOH once the shift happens?
3) What happens to the amount of $CaBr_2$ once the shift happens?
4) What happens to the amount of $Ca(OH)_2$ once the shift happens?

Jumpstart #1-G

$Ca(OH)_2 + 2KBr \leftrightarrow 2KOH + CaBr_2$

1) Add extra KBr

= shift right, need to use it up!

2) Amount of KOH

= increases, you moved toward it so you made some more

3) Amount of $CaBr_2$

= increases, you moved toward it so you made some more

4) Amount of $Ca(OH)_2$

= decreases, you moved away from it so you used some up

Page 237 – KCQ Target: I can pay attention to small details when doing Le Chatelier's principle problems

Things to look for BEFORE answering an equilibrium problem		
Stressor	**Question**	**What does it tell us?**
Increase or decrease [ ] products or reactants	**Which phase?**	• **Gas, aqueous - change things** • **Solid, Liquid – DON'T CHANGE ANYTHING!**
Increase or decrease T	**Endo or exo?**	• **Endo = absorbed, so it is a REACTANT** • **Exo = released, so it is a PRODUCT**
Increase or decrease total Pressure *(Same as change in volume or number of moles of gas)*	**How many moles of GAS are on each side of the equation?**	• **Increase pressure = move to side with FEWER moles of gas** • **Decrease pressure = move to side with MORE moles of gas**