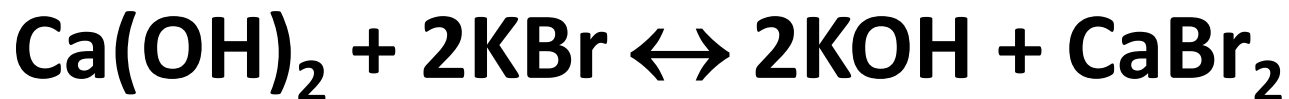
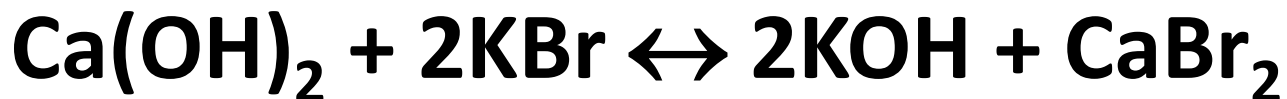


# Jumpstart #1-G



- 1) Which way does equilibrium shift if you add extra KBr
- 2) What happens to the amount of KOH once the shift happens?
- 3) What happens to the amount of  $\text{CaBr}_2$  once the shift happens?
- 4) What happens to the amount of  $\text{Ca(OH)}_2$  once the shift happens?

# Jumpstart #1-G



1) Add extra KBr

*= shift right, need to use it up!*

2) Amount of KOH

*= increases, you moved toward it so  
you made some more*

3) Amount of CaBr<sub>2</sub>

*= increases, you moved toward it so  
you made some more*

4) Amount of Ca(OH)<sub>2</sub>

*= decreases, you moved away from it  
so you used some up*

# Page 237 – KCQ Target: I can pay attention to small details when doing Le Chatelier's principle problems

## Things to look for BEFORE answering an equilibrium problem

Stressor	Question	What does it tell us?
Increase or decrease [ ] products or reactants	Which phase?	<ul style="list-style-type: none"><li>• Gas, aqueous - change things</li><li>• Solid, Liquid – DON'T CHANGE ANYTHING!</li></ul>
Increase or decrease T	Endo or exo?	<ul style="list-style-type: none"><li>• Endo = absorbed, so it is a REACTANT</li><li>• Exo = released, so it is a PRODUCT</li></ul>
Increase or decrease total Pressure <i>(Same as change in volume or number of moles of gas)</i>	How many moles of GAS are on each side of the equation?	<ul style="list-style-type: none"><li>• Increase pressure = move to side with FEWER moles of gas</li><li>• Decrease pressure = move to side with MORE moles of gas</li></ul>